

1. 0.10 g of hydrogen was formed when zinc reacts with sulfuric acid. How many moles of zinc are required?
2. How many grams of sodium sulfate will be formed when you start with 345 g of sodium hydroxide and excess sulfuric acid?
3. Lithium nitrate reacts with lead (IV) sulfate. How many moles of lithium nitrate are needed to make 250 g of lithium sulfate?
4. How many grams of copper (II) hydroxide can be prepared from 12.7 g of copper (II) nitrate and excess sodium hydroxide? How many moles of sodium hydroxide were used?
5. If 10.8 g of copper (II) sulfate is reacted with zinc how many moles of copper metal will be recovered from the solution?
6. 50.0 g of potassium chlorate are heated how many moles of oxygen are formed?
7. How many moles of oxygen are formed when 49.89 g of potassium chlorate decomposes?
8. When acetic acid reacts with aluminum hydroxide, water and aluminum acetate are formed. If I use 125 g of acetic acid how many moles of each of my products can I form? How many grams of aluminum hydroxide are used?
9. A mass of 21.5 grams of calcium hydroxide reacts with an excess of phosphoric acid. What is the mass of calcium phosphate could be recovered from solution?
10. When 13.5 grams of methane burn how many grams of each product are formed?