



4. A compound is 38.62 g Na, 26.9g S, and 53.8g O. Its formula mass is 284 g. What is its empirical and molecular formula?
5. A compound is 40% Carbon, 6.7% Hydrogen, and 53.3% Oxygen. Its formula mass is 180 g. What is its empirical and molecular formula?
6. A compound has 76.54% Carbon, 12.13% Hydrogen, and 11.33% Oxygen. The formula mass is 264.36 g. What is its empirical and molecular formula?