Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**SCH3U1 – Molar Concentration Note**

**Molarity-**

**Equation:**

**Example 1:**

Suppose we had 3.10 mole of sucrose and proceeded to mix it into exactly 1.25 liter of solution. What would be the molarity of this solution?

**Example 2:**

What is the molarity when 0.75 mol is dissolved in 2500 mL of solution?

**Example 3:**

Suppose you had 58.44 grams of NaCl and you dissolved it in exactly 2.00 L of solution. What would be the molarity of the solution?

**Example 4:**

Calculate the molarity of 25.0 grams of KBr dissolved in 750.0 mL.

**Example 5:**

80.0 grams of glucose (C6H12O6) is dissolved in enough water to make 1.00 L of solution. What is its molarity?